



## MILD, ASYMMETRIC DIELS-ALDER CYCLOADDITIONS OF ELECTRONICALLY MATCHED 2-PYRONES AND VINYL ETHERS

Gary H. Posner,\* Jean-Christophe Carry,<sup>1</sup> Jae Kyoo Lee, D. Scott Bull, and Haiyan Dai<sup>2</sup>

Department of Chemistry, School of Arts and Sciences,

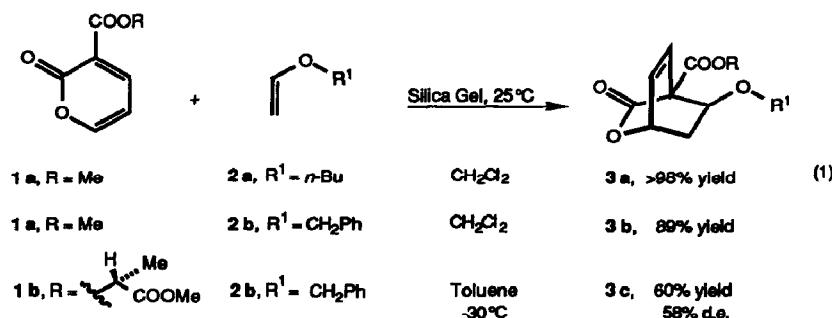
The Johns Hopkins University, Baltimore, MD 21218, USA

**Summary:** Silica gel and a TADDOL-complexed titanium IV Lewis acid are shown to promote mild, practical, asymmetric [4+2]-cycloadditions of electron-poor 2-pyrone-3-carboxylates with electron-rich vinyl ethers to form isolable and useful bicyclic lactone adducts.

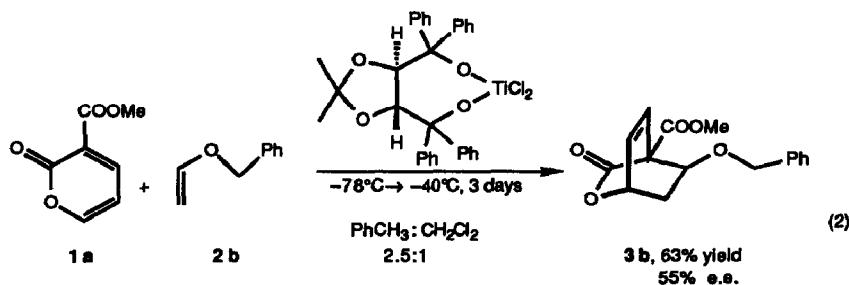
For almost a decade, we have been exploring use of electronically matched 2-pyrones and vinyl ethers for mild Diels-Alder cycloadditions leading to isolable and synthetically useful bicyclic lactone adducts.<sup>3</sup> Recently, others have begun to appreciate the synthetic potential of such inverse-electron-demand [4+2]-cycloadditions.<sup>4</sup> To control the absolute stereochemistry of such Diels-Alder cycloadditions, we have successfully used chiral auxiliaries attached to the vinyl ether dienophiles and separately to the pyrone dienes, as well as chiral auxiliaries bound to Lewis acidic lanthanide shift reagents.<sup>3</sup> To make these cycloadditions even more practical and operationally convenient, we have recently been exploring ways to use simple and readily available 2-pyrones and acid catalysts, and we now report some promising preliminary results.

Porous solids like silica,<sup>5</sup> alumina,<sup>6</sup> zeolites,<sup>7</sup> as well as montmorillonite clays<sup>8</sup> and molecular sieves<sup>9a</sup> have been used to facilitate many organic reactions, including some [4+2]-cycloadditions. These studies of [4+2]-cycloadditions promoted by porous solids have generally focused on model reactions involving cyclopentadiene and a variety of simple dienophiles. Montmorillonite clay-promoted reactions using (-)-menthyl acrylate and (R)-O-acryloyl pantolactone have also been studied.<sup>8b</sup> These dienophiles were shown to produce product *endo/exo* ratios of >9:1 with d.e. of 38% and 53%, respectively. In addition to these model studies, the synthetic utility of this methodology has been demonstrated in the preparation of ( $\pm$ )-verrucarol, with the key step involving an alumina-catalyzed intramolecular cycloaddition.<sup>6c</sup> We have now found that commercial chromatography silica gel (EM Science #60, 5 g per mmol of pyrone) promotes effective Diels-Alder cycloadditions of commercial 3-methoxycarbonyl-2-pyrone (1a) with butyl vinyl ether and with benzyl vinyl ether at room temperature to form

only the *endo* bicyclic diastereomer shown in eq. 1. Likewise, the enantiomerically pure pyrone lactate ester **1b**, prepared easily in three steps from pyrone methyl ester **1a**,<sup>3e</sup> reacted with benzyl vinyl ether in toluene over calcined (200°C, 16 hr.) Syloid 221 silica gel<sup>10</sup> to afford only *endo*-bicycloadducts (e.g. **3c**) as an approximately 4:1 mixture of diastereomers.<sup>11</sup> These [4+2]-cycloadditions facilitated by a very inexpensive, readily available porous solid are excellent examples of atom-economical processes<sup>12</sup> in which all of the atoms in both reactants appear in the product. Also, these heterogeneous Diels-Alder processes are environmentally benign, generating no harmful by-products. In the absence of silica gel, no cycloaddition occurred.<sup>13</sup>



Lewis acids like titanium IV complexed with homochiral ligands have been used successfully in a variety of asymmetric reactions.<sup>9,14</sup> As a logical extension of our recent studies with homochiral lanthanide shift reagents,<sup>3e</sup> we have now found that commercial 3-methoxycarbonyl-2-pyrone (**1a**) and the tartrate-derived TADDOL-complexed titanium IV species<sup>9</sup> shown in eq. 2 react with benzyl vinyl ether under very mild conditions to produce only *endo*-bicycloadduct **3b** as a single diastereomer in 55% enantiomeric purity. The enantiomeric purity of bicyclic lactone **3b** formed in eq. 2 was determined using an enantiopure NMR shift reagent in comparison to the same NMR determination using independently prepared racemic bicyclic lactone **3b**.



In summary, the porous solid-promoted cycloadditions shown in eq. 1 and the homochiral Lewis acid-promoted cycloaddition shown in eq. 2 represent important initial progress in developing practical, simple, and environmentally benign asymmetric syntheses of versatile bicyclic lactones under reaction conditions mild enough to prevent loss of CO<sub>2</sub> from the initial bicycloadducts.<sup>3a,13</sup> We are continuing to explore ways to maximize operational simplicity and stereocontrol.

### Acknowledgment

We thank the donors of The Petroleum Research Fund, administered by the ACS, for support of this research.

### References

1. Current address: Rhône-Poulenc Rorer, Vitry sur Seine, France.
2. Current address: The Procter and Gamble Co., Miami Valley Labs, Cincinnati, Ohio.
3. a) Posner, G.H.; Harrison, W. *J. Chem. Soc. Chem. Commun.*, **1985**, 1786; b) Posner, G.H.; Wetlaufer, D.G. *Tetrahedron Lett.*, **1986**, *27*, 667; c) Posner, G.H.; Wetlaufer, D.G. *J. Am. Chem. Soc.*, **1986**, *108*, 7373; d) Posner, G.H.; Switzer, C. *J. Org. Chem.*, **1987**, *52*, 1642; e) Posner, G.H.; Haces, A.; Harrison, W.; Kinter, C. *ibid.*, **1987**, *52*, 4836; d) Posner, G.H.; Nelson, T.D. *Tetrahedron*, **1990**, *46*, 4573; e) Posner, G.H.; Kinter, C.M. *J. Org. Chem.*, **1990**, *55*, 3967; f) Posner, G.H.; Nelson, T.D. *ibid.*, **1991**, *56*, 4339; g) Posner, G.H.; Nelson, T.D.; Afarinkia, K. *Tetrahedron Lett.*, **1991**, *32*, 5295; h) Posner, G.H.; Nelson, T.D.; Kinter, C.M.; Johnson, N. *J. Org. Chem.*, **1992**, *57*, 4083; i) Posner, G.H.; Vinader, V.; Afarinkia, K. *ibid.*, **1992**, *57*, 4088; j) Posner, G.H.; Nelson, T.D.; Guyton, K.Z.; Kensler, T.W. *J. Med. Chem.*, **1992**, *35*, 3280; k) Afarinkia, K.; Posner, G.H. *Tetrahedron Lett.*, **1992**, *33*, 7839; l) Posner, G.H.; Carry, J.-C.; Anjeh, T.E.N.; French, A.N. *J. Org. Chem.*, **1992**, *57*, 7012; m) Posner, G.H.; Dai, H. *BioMed. Chem. Lett.*, **1993**, *3*, 1829; n) Posner, G.H.; Dai, H.; Afarinkia, K.; Murthy, N.N.; Guyton, K.Z.; Kensler, T.W. *J. Org. Chem.*, **1993**, *58*, 7209; o) For an extensive review, see Afarinkia, K.; Vinader, V.; Nelson, T.D.; Posner, G.H. *Tetrahedron*, **1992**, *48*, 9111-9171; p) For a short review, see Kvita, V.; Fischer, W. *Chimia*, **1993**, *47*, 3-18.
4. a) Swarbrick, T.M.; Marko, I.E.; Kennard, L. *Tetrahedron Lett.*, **1991**, *32*, 2549; b) Marko, I.E.; Seres, P.; Swarbrick, T.M.; Staton, I.; Adams, H. *ibid.*, **1992**, *33*, 5649; c) Marko, I.E.; Seres, P.; Evans, G.R.; Swarbrick, T.M. *ibid.*, **1993**, *34*, 7305; d) Marko, I.E.; Evans, G.R. *ibid.*, **1993**, *34*, 7309.
5. Veselovsky, V.V.; Gybin, A.S.; Lozanova, A.V.; Moiseenkov, A.M.; Smith, W.A.; Caple, R. *Tetrahedron Lett.*, **1988**, *29*, 175-178.
6. a) Parlan, H.; Baumann, R. *Angew. Chem. Int. Ed. Engl.*, **1981**, *20*, 1014; b) Kabalka, G.W.; Pagni, R.M.; Bains, S.; Hondrogiannis, G.; Plesco, M.; Kurt, R.; Cox, D.; Green, J. *Tetrahedron: Asymmetry*, **1991**, *2*, 1283-1294; c) Koreeda, M.; Ricca, D.J.; Loengo, J. *I. J. Org. Chem.*, **1988**, *53*, 5585-5588.
7. a) Dessau, R.M. *J. Chem. Soc. Chem. Commun.*, **1986**, 1167-1168; b) Ipaktschi, J., Z. *Naturforsch.*, **1986**, *416*, 496.

8. a) Laszlo, P. *Accts. Chem. Res.*, **1986**, *19*, 121-127; b) Cativiela, C.; Figueras, F.; Fraile, J.M.; Garcia, J.I.; Mayoral, J.A. *Tetrahedron: Asymmetry*, **1993**, *4*, 223-228, and references therein.
9. a) Narasaka, K.; Iwasara, N.; Inoue, M.; Yamada, T.; Nakashima, M.; Sugimori, J. *J. Am. Chem. Soc.*, **1989**, *111*, 5340-5345; b) Hayashi, Y.; Narasaka, K. *Chem. Lett.*, **1990**, 1295-1298; c) Narasaka, K.; Kusama, H.; Hagashi, Y. *Bull. Chem. Soc. Japan*, **1991**, *64*, 1471-1478; d) Narasaka, K.; Inoue, M.; Okada, N. *Chem. Lett.*, **1986**, 1109; e) Narasaka, K.; Inoue, M.; Yamada, T.; Sugimon, J.; Zwasawa, N. *ibid.*; f) Narasaka, K.; Inoue, M.; Yamada, T. *Chem. Lett.*, **1986**, 1967; g) Corey, E.J.; Matsumura, Y. *Tetrahedron Lett.*, **1991**, 6289.
10. 160 Å° Pore diameter; 400 m<sup>2</sup>/gm surface area; 7 μ particle size; cf. Cativiela, C.; Figueras, F.; Garcia, J.I.; Mayoral, J.A.; Pires, E.; Royo, A.J. *Tetrahedron: Asymmetry*, **1993**, *4*, 621-624, and references therein.
11. All new compounds were characterized spectroscopically and by combustion analysis or by <sup>13</sup>C NMR spectroscopy combined with HRMS.
12. a) Trost, B.M. *Science*, **1991**, *254*, 1471-1477; b) Trost, B.M.; Zhi, L. *Tetrahedron Lett.*, **1992**, *33*, 1831-1834.
13. For use of higher temperatures and some Lewis acids to promote closely related cycloadditions of 2-pyrone leading ultimately to aromatic products, see Boger, D.L.; Mullican, M.D. *J. Org. Chem.*, **1984**, *49*, 4033.
14. Engler, T.A.; Letavic, M.A.; Reddy, J.P. *J. Am. Chem. Soc.*, **1991**, *113*, 5068-5070; b) Takashi, H.; Kawakita, T.; Yoshioka, M.; Kobayashi, S.; Ohno, M. *Tetrahedron Lett.*, **1989**, *30*, 7095-7098; c) Terada, M.; Mikani, K.; Nakai, T. *ibid.*, **1991**, *32*, 935-938. For other homochiral Lewis acids, see d) Denmark, S.E.; Jacobs, R.T.; Dai-Ho, G.; Wilson, S. *Organometallics*, **1990**, *9*, 3015-3019; e) Kobayashi, S.; Murakami, M.; Harada, T.; Mukaiyama, T. *Chem. Lett.*, **1991**, 1341-1344; f) Hawkins, J.M.; Loren, S. *J. Am. Chem. Soc.*, **1991**, *113*, 7794-7795; g) Sarton, D.; Saffrich, J.; Helmchen, G. *Synlett*, **1990**, 197-198; h) Corey, E.J.; Loh, T.-P.; Roper, T.D.; Azimioara, M.D.; Noe, M.C. *J. Am. Chem. Soc.*, **1992**, *114*, 8290-8292 and references therein; i) Schmidt, B.; Seebach, D. *Angew. Chem. Int. Ed. Engl.*, **1991**, *30*, 1321; j) Seebach, D.; Behrendt, L.; Felix, O. *ibid.*, **1991**, *30*, 1008; k) Schmidt, B.; Seebach, D. *ibid.*, **1991**, *30*, 99.

(Received in USA 19 November 1993; revised 14 December 1993; accepted 17 December 1993)